PURAMUFTI PUBLIC SCHOOL & COLLEGE

PRAYAGRAJ, PURAMUFTI

CLASS – 12 (E.M+H.M.)

SUBJECT:- HINDI

SUMMER VACATION HOME-WORK (2026-27)

हिन्दी गद्य साहित्य

1. कैन्हया लाल मिश्र प्रभाकर का चित्र बनाकर उनका जीवन परिचय देते हुए साहित्यिक परिचय और उनकी रचनाएँ लिखिए। (पेज नं. 53)
2. गद्यांश पर आधारित प्रश्न के उत्तर दीजिए।

   संकेत- नश्तर तेज था---------- पर कितनी अजेय।

   (अ) उपर्युक्त गद्यांश के लेखक एवं पाठ का नाम लिखिए।

   (ब) मनुष्य- मनुष्य के बीच दीवारे किसने खड़ी की?

   (स) लेखक के प्रश्न के उत्तर में मदर टेरेसा ने क्या कहा?

   (द) प्रस्तुत गद्यांश में लेखक ने किस बात की ओर संकेत किया?

हिन्दी पद्य साहित्य

1. शुक्ल जी द्वारा रचित काल विभाजन का सचित्र वर्णन कीजिए तथा उनकी समय सीमा का निर्धारण कीजिए। उदाहरण:- आदिकाल, पूर्वमध्य काल, उत्तर मध्यकाल, और आधुनिक काल।
2. मैथिलीशरण गुप्त का जीवन परिचय, साहित्यिक परिचय तथा कृतियों का वर्णन कीजिए।

कथा साहित्य

1. फणीश्वर नाथ 'रेणु' द्वारा रचित पंचलाइट कहानी की कथावस्तु अपने शब्दों में लिखिए। (पेज नं.-261)
2. अमरकान्त द्वारा रचित बहादुर कहानी की समीक्षा कीजिए। (पेज नं.- 266)

खण्डकाव्य

1. आलोक वृत खण्डकाव्य के आधार पर महात्मा गांधी का चित्र बनाकर प्रथम, द्वितीय और तृतीय सर्ग की कथावस्तु अपने शब्दों मे लिखिए। (पेज नं.- 314)

संस्कृत दिगदर्शिका

1. 'भोजस्यौदार्यम' पाठ का हिन्दी में अनुवाद कीजिए।

व्याकरण

1. रस की परिभाषा देते हुए सभी रसो का वर्णन चार्ट पेपर पर कीजिए। (पेज नं.-245)
2. अलंकार की परिभाषा दीजिए तथा अनुप्रास, यमक, श्लेष, उत्प्रेक्षा तथा भ्रान्तिमान अलंकार का लक्षण एवं उदाहरण सहित लिखिए। (पेज नं.-255)

SUMMER VACATION ASSIGNMENT

SUBJECT – MATHS (Class 12th-E.M.)

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Question 1: Write in detail about the life of any Indian mathematician (**Aryabhatta)** and tell about his great achievements in the field of mathematics and also paste a coloured picture of the concerned mathematician.

Question 2:

LESSON	EXERCISE (QUESTION NO.)
Inverse trigonometric functions	miscellaneous questionnaire
Matrix	Exercise 3.2 (6,7,8,9,10)
	Exercise 3.3 (2,3,7,8,9)

Question 3: If $A = \begin{bmatrix} 1 & 2 & 3 \\ 3 & -2 & 1 \\ 4 & 2 & 1 \end{bmatrix}$ then prove $A^3 - 23A - 40I = 0$.

नोट:- सभी प्रश्न फ□ल क□ प□ज□ में लिखन/बन ना□ है तथ फ□ल की सज□ट अति आवश्यक है|

12th Summer Home Work Assignment

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PRAYAGRAJ

SUBJECT – ENGLISH (Class 12th)

Make a Project File

Topics –

1- Narration (Direct & Indirect Speech)

Change the following sentences into Indirect Speech –

- Rahul said, "I am reading a book."
- She said, "I have completed my homework."
- The teacher said, "Work hard for success."
- Mohan said, "I can solve this question."
- Father said, "Do not waste your time."
- Riya said, "We are going to Delhi tomorrow."
- He said, "I was playing cricket."
- Mother said, "Please help me in the kitchen."
- The boy said, "I shall win the match."
- Sita said, "I may come late today."
- The teacher said, "Honesty is the best policy."
- He said to me, "Are you feeling well?"
- She said, "What a beautiful flower it is!"
- Father said to his son, "Always speak the truth."
- Ravi said, "I did not attend the class yesterday."
- The stranger said, "Can you tell me the way to the

station?”

- The doctor said, “Take this medicine regularly.”
- My friend said, “I will help you.”
- She said to me, “Why are you late?”
- The policeman said, “Do not cross the road here.”

2- Transformation of Sentences

A- Change into Negative Sentences

- Everyone likes honesty.
- He always comes on time.
- She is very intelligent.
- I remember his name.
- They accepted the proposal.
- The boy was happy.
- We obey our teachers.
- He completed his work.
- She sings beautifully.
- India won the match.

B- Change into Interrogative Sentences

- He is your friend.
- She plays badminton well.
- They were present in the class.
- You know the answer.
- He can speak English fluently.
- She has completed the work.
- They enjoyed the picnic.
- The train arrived on time.
- We should help the poor.

- You like reading books.

3- Letter Writing

Formal Letters

- Write a letter to the Principal requesting extra classes for English.
- Write a letter to the Editor about increasing pollution in your city.
- Write a letter to the Municipal Commissioner complaining about dirty roads and garbage in your locality.
- Write a letter to the Principal requesting leave for three days due to illness.
- Write a letter to the Editor about the importance of road safety rules.

Informal Letters

- Write a letter to your friend describing how you spent your summer vacation.
- Write a letter to your younger brother advising him to study regularly.
- Write a letter to your friend inviting him/her to your birthday party.
- Write a letter to your friend describing your school annual function.
- Write a letter to your cousin describing the importance of physical exercise.

4- Article Writing

Write an article on any one topic –

- Importance of Education

- Social Media and Students
- Importance of Discipline
- Save Environment Save Earth
- Role of Technology in Education

5- Notice Writing

- Prepare a notice on any one topic –
- Debate Competition
- Independence Day Celebration
- Lost and Found
- Sports Trial
- Cleanliness Campaign

Literature Section

6- Write Summary of the following chapters –

- The Last Lesson
- Lost Spring
- The Third Level
- My Mother at Sixty-Six

7- Write Character Sketch of –

- Franz
- M. Hamel
- Saheb
- Mukesh

8- Write the Central Idea / Theme of all chapters and poem.

9- Literature Activity

Draw/Paste pictures related to the chapters and write five lines about each.

10- Creative Activity

Make a Mind Map or Flow Chart of any one chapter.

11- Newspaper Activity

Pick out any five difficult words from the newspaper everyday. Write their meanings and make sentences. (Only 20 words)

12- Quote Writing Activity

Write any 10 inspirational quotations related to education, mother, language or life.

13- Grammar Chart Activity

Prepare colorful charts on Tenses, Narration Rules and Transformation Rules.

14- Reading Activity

Read any English newspaper or story book for 15 minutes daily and write a short report about what you learned.

15- Project Decoration

Use colored headings, borders, sketches and pictures to make the project attractive.

By Suryansh Pandey Sir

PURAMUFTI PUBLIC SCHOOL

CLASS 12 | CHEMISTRY

DPP | **CHAPTER: SOLUTIONS** | **BOARD LEVEL HARD**

NUMERICAL QUESTIONS – ANSWER ONLY

1. The vapour pressure of pure solvent A is 720 mmHg at 25°C and that of its solution containing a non-volatile solute is 684 mmHg. What is the mole fraction of the solute in the solution?
0.050

2. A solution is prepared by dissolving 4.34 g of a non-volatile solute in 100 g of water. The freezing point of solution is −0.372°C. Calculate the molar mass of solute. (K_f for water = 1.86 K kg mol^{-1})
186

3. The boiling point of a solution containing 3.42 g of a non-volatile solute in 50 g of water is 100.26°C. Calculate the molar mass of the solute. (K_b for water = 0.52 K kg mol^{-1})
180

4. Osmotic pressure of a solution is 2.47 atm at 27°C. If 1.00 g of a non-electrolyte is dissolved in 250 mL of the solution, find its molar mass.
42

5. A 0.05 M solution of a non-electrolyte has an osmotic pressure of 1.23 atm at 25°C. Find the molar mass of the solute.
82

6. The freezing point of 0.10 m solution of an electrolyte is −0.558°C. Calculate the van 't Hoff factor (i). (K_f for water = 1.86 K kg mol^{-1})
3

7. A solution of glucose (molar mass = 180 g mol^{-1}) exerts an osmotic pressure of 4.11 atm at 27°C. What is the molarity of the solution?
1.00

8. The relative lowering of vapour pressure for a solution is 0.08. If the solution contains 18 g of solute, find the mass of solvent.
207 g

9. A solution contains 10 g of urea (molar mass = 60 g mol^{-1}) in 90 g of water. Calculate the boiling point of the solution. (K_b for water = 0.52 K kg mol^{-1})
100.57°C

10. The osmotic pressure of a solution containing 0.5 g of a non-electrolyte in 100 mL is 0.615 atm at 30°C. Find the molar mass of the solute.
41

11. 2.56 g of a non-volatile nonelectrolyte is dissolved in 80 g of benzene. The boiling point of solution is 80.26°C. Calculate the molar mass of the solute. (K_b for benzene = 2.53 K kg mol^{-1})
128

12. The vapour pressure of water at 25°C is 23.8 mmHg. If a solution has vapour pressure of 22.61 mmHg, find mole fraction of solute.
0.050

13. The freezing point of 0.01 m solution of an electrolyte is −0.0558°C. Calculate the van 't Hoff factor (i). (K_f for water = 1.86 K kg mol^{-1})
3

14. 0.25 g of a non-electrolyte is dissolved in 50 mL of solution. The osmotic pressure is 0.246 atm at 27°C. Find the molar mass of the solute.
41

15. A 1.0 m glucose solution boils at 100.52°C. What will be the boiling point of 0.5 m glucose solution? (K_b for water = 0.52 K kg mol^{-1})
100.26°C

16. The osmotic pressure of 0.01 M solution of an electrolyte is 0.513 atm at 27°C. Calculate the van 't Hoff factor (i).
2

17. A solution containing 1.8 g of glucose in 100 g of water freezes at −0.186°C. Find molar mass of glucose. (K_f for water = 1.86 K kg mol^{-1})
180

18. The boiling point elevation for 0.02 m solution of an electrolyte is 0.052°C. Calculate the van 't Hoff factor (i). (K_b for water = 0.52 K kg mol^{-1})
5

21. A 0.10 m solution of an electrolyte has freezing point −0.372°C. Calculate i. (K_f for water = 1.86 K kg mol^{-1})
2

22. The osmotic pressure of a solution is 1.538 atm at 25°C. If the molar mass of solute is 154 g mol^{-1}, find molarity.
0.200

23. The boiling point of 0.01 m solution of an electrolyte is 100.026°C. Calculate i. (K_b for water = 0.52 K kg mol^{-1})
5

24. A solution containing 34.2 g of sucrose (molar mass = 342 g mol^{-1}) in 200 g of water has depression in freezing point 0.372°C. Find i.
1

25. The vapour pressure of a solution is 690 mmHg. The vapour pressure of pure solvent is 750 mmHg. Find mole fraction of solute.
0.080

26. The osmotic pressure of 0.02 M solution of an electrolyte is 1.026 atm at 27°C. Calculate i.
2

27. 3.0 g of an electrolyte is dissolved in 100 g of water. The boiling point of solution is 100.156°C. Calculate molar mass of electrolyte. (K_b for water = 0.52 K kg mol^{-1}, i = 2)
144

28. The freezing point of 0.02 m solution of an electrolyte is −0.2232°C. Calculate i.
6

29. The osmotic pressure of 0.05 M solution of a non-electrolyte is 1.269 atm at 27°C. Find molar mass of solute.
82

30. A solution of an electrolyte shows 20% lowering in vapour pressure. If i = 2, find mole fraction of solute.
0.111

31. The boiling point elevation of a 0.02 m solution of an electrolyte is 0.0104°C. Calculate i.
1

32. 1.64 g of a non-electrolyte in 100 mL solution exerts osmotic pressure 0.492 atm at 25°C. Find molar mass.
84

33. The freezing point of 0.1 m solution of an electrolyte is −0.558°C. Calculate i.
3

34. A solution contains 0.1 mole of non-volatile solute in 500 g of solvent. Relative lowering of vapour pressure is 0.02. Molality of solution is:
0.2

35. The osmotic pressure of 0.005 M solution of an electrolyte is 0.2565 atm at 27°C. Calculate i.
5

36. 5.0 g of urea is dissolved in 95 g of water. Calculate the boiling point of the solution.
100.27°C

37. A solution has mole fraction of solute 0.04. If the vapour pressure of solvent is 17.4 mmHg, what is vapour pressure of solution?
16.70 mmHg

38. The osmotic pressure of 0.1 M solution of an electrolyte is 2.568 atm at 27°C. Calculate i.
2

39. The freezing point depression of 0.05 m solution of an electrolyte is 0.279°C. Calculate i.
3

40. 2.5 g of a non-electrolye is dissolved in 75 g of 07.5 g of benzene pressure of solution is. (K_f for benzene:
12.0

41. The vapour pressure of solution is 23.1 mmHg. The vapour pressure of solvent is 24.3 mmHg. Find mole fraction of solute.
0.049

42. Osmotic pressure of a solution is 0.820 atm at 20°C. If molar mass of solute is 205 g mol^{-1}, find molarity.
0.100

43. The boiling point elevation of 0.01 m solution of electrolyte is 0.026°C. Calculate i.
5

44. A solution containing 10 g of glucose in 250 g of water has depression in freezing point 0.186°C. Find molar mass of glucose.
180

45. The vapour pressure of solvent is 600 mmHg. The relative lowering of vapour pressure of solution is 0.06. Find vapour pressure of solution.
564 mmHg

46. The osmotic pressure of 0.1 M solution of an electrolyte is 3.0816 atm at 27°C. Calculate i.
3

47. 4.0 g of an electrolyte (i = 2) is dissolved in 200 g of water. Calculate the boiling point of solution.
100.21°C

48. The freezing point of 0.02 m solution of an electrolyte is −0.372°C. Calculate i.
10

49. A 0.02 M solution of an electrolyte has osmotic pressure 1.025 atm at 25°C. Calculate i.
5

50. 1.0 g of non-electrolyte in 50 mL solution exerts osmotic pressure 0.123 atm at 25°C. Find molar mass.
82

51. The vapour pressure of solution is 708 mmHg. The vapour pressure of solvent is 750 mmHg. Find mole fraction of solute.
0.056

52. The boiling point elevation of 0.05 m solution of an electrolyte is 0.065°C. Calculate i.
2.5

53. The freezing point of 0.01 m solution of an electrolyte is −0.1116°C. Calculate i.
6

54. Osmotic pressure of a solution is 0.615 atm at 27°C. If 0.75 g solute is dissolved in 100 mL of solution, find molar mass.
41

55. The vapour pressure of solution is 22.8 mmHg. The vapour pressure of solvent is 24 mmHg. Find mole fraction of solvent.
0.950

56. 3.42 g of urea is dissolved in 100 g of water. Calculate osmotic pressure at 27°C.
1.44 atm

57. The boiling point of 0.01 m solution of electrolyte is 100.052°C. Calculate i.
10

58. The freezing point depression of 0.1 m solution of an electrolyte is 0.558°C. Calculate i.
3

59. A solution has osmotic pressure 1.538 atm at 25°C. If molarity is 0.2, calculate i.
2

60. The boiling point of a solution is 100.78°C. The molar elevation constant of solvent is 0.26 K kg mol^{-1}. Molality of solution is:
3.0 m